

Oxidation reaction path	Reaction rate coefficient (L moles ⁻¹ s ⁻¹) at 298 K	Temperature dependence $\frac{-E}{R}$ (K)
$O_3(aq) + SO_2 * H_2O \rightarrow S^{VI}$	$k_0 = 2.4 \times 10^4$	0
$O_3(aq) + HSO_3^- \rightarrow S^{VI}$	$k_1 = 3.5 \times 10^5$	-5530
$O_3(aq) + SO_3^{2-} \rightarrow S^{VI}$	$k_2 = 1.5 \times 10^9$	-5280
$H_2O_2(aq) + HSO_3^- \rightarrow S^{VI}$	$k_3 = 7.45 \times 10^7$	-4430